

Periodic Trends Worksheet 2

Name: Key
 Date:
 Block:

1. Arrange the following from largest atomic radius to smallest atomic radius:
 a) Po, O, Se, S



- b) W, Hg, Cs, La



2. Arrange the following from highest ionization energy to lowest ionization energy:
 a) At, F, Br, Cl



- b) Sb, Rb, Ag, I



3. Arrange the following from most electronegative to least electronegative:
 a) Bi, N, As, P, Sb



- b) Sc, Ge, Cu, Ti, Kr



4. Use the following particles to answer the questions below:

has more shells than others ∴ biggest.

Chemical Name:	8p 18e ⁻ Argon atom	21p 18e ⁻ Scandium ion	16p 16e ⁻ Sulfur atom	20p 20e ⁻ Calcium atom	19p 18e ⁻ Potassium ion
Chemical Formula:	Ar	Sc ³⁺	S	Ca	K ⁺

- a. Rank the particles in atomic size from smallest to biggest (you must use **chemical formulas** and *not* the chemical name):



- b. Which of the particles would have the greatest electronegativity?



- c. Which of the particles would have the greatest ionization energy?

